

Unit 19 Period 3 Elements & Their Oxides (Paper 1 & 3)

19.1 Properties of Period 3 elements and their oxides

The reactions of Na and Mg with water.

The trends in the reactions of the elements Na, Mg, Al, Si, P and S with oxygen, limited to the formation of Na_2O , MgO , Al_2O_3 , SiO_2 , P_4O_{10} , SO_2 and SO_3




The trend in the melting point of the highest oxides of the elements Na–S

The reactions of the oxides of the elements Na–S with water, limited to Na_2O , MgO , Al_2O_3 , SiO_2 , P_4O_{10} , SO_2 and SO_3 , and the pH of the solutions formed.

The structures of the acids and the anions formed when P_4O_{10} , SO_2 and SO_3 react with water.

You should be able to:

- explain the trend in the melting point of the oxides of the elements Na–S in terms of their structure and bonding
- explain the trends in the reactions of the oxides with water in terms of the type of bonding present in each oxide
- write equations for the reactions that occur between the oxides of the elements Na–S and given acids and bases.

		
Revision Done?	YES	NO